

## DISPLACEMENT REACTIONS OF HALOGENS

Paper: 2CR

Question: 4(b)

### Question

(b) Chlorine gas is bubbled into a colourless solution of potassium bromide.

Explain why the solution turns orange.

(2)

### Mark Scheme

(b)	Ex	An explanation linking the following two points	
		<b>M1</b> bromine / Br <sub>2</sub> is formed / displaced / produced	<b>REJECT</b> bromide for bromine
			<b>ACCEPT</b> bromine/Br <sub>2</sub> shown as the product in an equation
			<b>IGNORE</b> state of bromine
		<b>M2</b> as chlorine is more reactive (than bromine)	<b>REJECT</b> bromide/chloride